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368 Chapter 11 • Stoichiometry

Section 11.1.1 Objectives Describe the types of relationships indicated by a balanced chemical equation. State the mole ratios from a balanced chemical equation. Review Vocabulary reactant: the starting substance in a chemical reaction New Vocabulary stoichiometry mole ratio Defining Stoichiometry

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of the quantitative relationships

between the amounts of reactants used and the amounts of products formed by a chemical reaction

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and Change • Chapter 11 209
Stoichiometry Stoichiometry CHAPTER
11 SOLUTIONS MANUAL Section
11.1 Defining Stoichiometry pages
368–372 Practice Problems pages
371–372 1. Interpret the following
balanced chemical equations in terms
of particles, moles, and mass. Show
that the law of conservation of mass is

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Stoichiometry. stoichiometry. mole
ratio. excess reactant. limiting
reactant. The study of quantitative
relationships between the amounts
of.... In a balanced equation, the ratio
between the numbers of moles.... A
reactant that remains after a chemical
reaction stops.

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following: How much oxygen is needed to ensure complete combustion of a ...

Chapter 11.4: Stoichiometry -
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Chapter 11 Stoichiometry Study Guide 11.1 Defining Stoichiometry MAIN Idea The amount of each reactant present at the start of a chemical reaction

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determines how much product can form. 11.2 Stoichiometric Calculations
MAIN Idea The solution to every stoichiometric problem requires a balanced chemical equation. 11.3 Limiting Reactants
MAIN Idea A chemical reaction Chapter 11: Stoichiometry
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These topics are followed by Key discussions on writing names and formulas of compounds; chemical formulas and the mole; chemical reactions; calculations based on equations; gases; and the properties of a liquid. The remaining chapters examine the solutions; acids; bases; salts; oxidation-reduction reactions; electrochemistry; chemical kinetics and equilibrium; and nuclear, organic, and biological chemistry. This study guide will be of great value to chemistry teachers and students.

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